

Additions and Corrections

1972, Volume 11

John A. Ferguson, Thomas J. Meyer*, and David G. Whitten*: Metalloporphyrin Redox Chemistry. Unusual Electron-Transfer Behavior in the Oxidation of Lead Porphyrins.

Page 2770. In column 1, lines 12 and 13, the g values from epr measurements should read 2.0158 and 2.0161, respectively.—T. J. Meyer

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Frank J. Weigert and John D. Roberts*: Nuclear Magnetic Resonance Spectroscopy. Spin-Spin Coupling of Carbon to Phosphorus, Mercury, Nitrogen, and Other Elements.

Page 315. We are indebted to Professor W. T. Ford for pointing out that the data listed for tri-*n*-propylboron in Table III are incorrect. The broad line corresponding to the α carbon is actually at 31.7 ppm, the β carbon resonance comes at 18.4 ppm, and one from the γ carbon at 17.8 ppm.—J. D. Roberts

R. J. West and S. F. Lincoln*: Exchange of Acetonitrile on Complexes of Nickel(II) and Cobalt(II) Formed with 2,2',2''-Triamino-triethylamine and 2,2',2''-Tri(*N,N*-dimethylamino)triethylamine.

Page 495. Equation 3 should read:

$$T_{2p}^{-1} = \left(\frac{\tau_M}{P_M} \left[\frac{(T_{2M}^{-1} + \tau_M^{-1})^2 + \Delta\omega_M^2}{T_{2M}^{-2} + (T_{2M}\tau_M)^{-1} + \Delta\omega_M^2} \right] \right)^{-1} + \left(\frac{T_{2M}'}{P_M'} \right)^{-1}$$

Page 496. The left-hand ordinate of Figure 2 should read: $\ln(T_{2p}m)$.—S. F. Lincoln

G. M. Bancroft*, H. C. Clark, R. G. Kidd, A. T. Rake, and H. G. Spinney: Manganese-55 Nuclear Magnetic Resonance Studies of Compounds Containing Mn-M (M = Si, Ge, Sn, Pb) Bonds and the Signs of e^2qQ in Manganese(I) Compounds.

Page 730. In Table III, the observed e^2qQ for $Mn(CO)_5SnCl_3$ should read 7.3 *not* 73.—G. M. Bancroft

J. Pradilla-Sorzano and John P. Fackler, Jr.*: Base Adducts of β -Ketoenolates. V. Crystal and Molecular Structures of *cis*-Bis-(1,1,1,5,5,5-hexafluoro-2,4-pentanedionato)bis(pyridine)zinc(II) and -copper(II).

Pages 1174 and 1175. In the title and several other places the

title compounds are incorrectly referred to as *cis*-bis(1,1,1,6,6,6-hexafluoro-2,4-pentanedionato)bis(pyridine)zinc(II) and -copper(II). The correct names are *cis*-bis(1,1,1,5,5,5-hexafluoro-2,4-pentanedionato)bis(pyridine)zinc(II) and -copper(II).—John P. Fackler, Jr.

J. Pradilla-Sorzano and John P. Fackler, Jr.*: Base Adducts of β -Ketoenolates. VI. Single-Crystal Electron Paramagnetic Resonance and Optical Studies of Copper(II)-Doped *cis*-Bis(hexafluoroacetylacetonato)bis(pyridine)zinc(II), Cu-Zn(F_6acac)₂(py)₂.

Page 1188. Equation 8 should read:

$$g = \cos^2 \left(\frac{\phi_2}{2} \right) g_{x^2-y^2} + \sin^2 \left(\frac{\phi_2}{2} \right) g_{3z^2-r^2}$$

R. Morassi, F. Mani, and L. Sacconi*: Spin Isomerism in Five-Coordinate Cobalt(II) Complexes of the Terdentate Ligands *N-R-N*-2-(diphenylphosphino)ethyl-*N,N'*-diethylethylenediamine (Donor Set NNP).

Page 1249. Reference 15 should read: R. Morassi and L. Sacconi, *J. Chem. Soc. A*, 1487 (1971).—L. Sacconi

G. B. Robertson and P. O. Whimp*: Crystal and Molecular Structure of a Stable π -Oxopropenylmanganese Tricarbonyl Complex.

Page 1741. In the Data Collection section, the formula for the standard deviation of an individual reflection (line 21) should read

$$\sigma(F_o) = [(LP)^{-2}\Delta I^2 + (\rho|F_o|^2)^2]^{1/2}/2|F_o|$$

while on line 34, the formula for $\sigma_s(F_o)$ should read

$$\sigma_s(F_o) = (LP)^{-1}\Delta I/2|F_o|$$

P. O. Whimp

Earl L. Muetterties: Stereochemical Lability of Eight-Coordinate Complexes.

Page 1964. The second sentence should read “. . . ; an X-ray analysis of the related [(*n*-C₄H₉)₄N]₃Mo(CN)₈ salt has shown that the Mo(CN)₈³⁻ ion. . .”.

Page 1966. In Table I, column entries for ¹³C nmr data were inadvertently transposed. The ¹³C δ_a and δ_c values for Zr[S₂CN(CH₂CH₂CH₃)₂]₄ are -180.9 and -140.7, respectively. The ¹³C δ_a and δ_c values for Ti[S₂CN(CH₂CH₂CH₃)₂]₄ are -148.5 and -107.0, respectively.

In the third paragraph the third sentence should read “The electronic spectra of [(*n*-C₄H₉)₄N]₄W(CN)₈. . .”. The fifth sentence should read “There was, however, no correspondence between the electronic spectrum of the propylammonium salt. . .”.—E. L. Muetterties